

Test Unit 3 SCH3U #10

Name: _____

 K T C A Assessment of Learning (Product)
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Multiple Choice: Circle the most correct answer. (10K)

1. What is the molar mass of AuCl_3 ?

- A) 96 g B) 130 g C) 232.5 g D) 303.6 g

2. Avogadro's number of representative particles is equal to one _____.

- A) kilogram B) gram C) Kelvin D) mole

3. How many moles of tungsten atoms are in 4.8×10^{22} atoms of tungsten?

- A) 8.0×10^2 moles B) 8.0×10^1 moles C) 1.3×10^{-1} moles D) 1.3×10^{-2} moles

4. Determine the mass in grams of 0.0489 mol cobalt.

- A) 0.00812 g B) 0.294 g C) 2.88 g D) 0.000830 g

5. What is the mass in grams of 1.02×10^{24} atoms manganese (Mn)?

- A) 0.112×10^1 g B) 0.169×10^6 g C) 9.30×10^{-1} g D) 9.30×10^6 g

6. Which of the following is/are true?

- i. Samples of water coming from different sources will still have the same percentage composition
- ii. The percentage composition of a compound can ONLY be calculated with knowledge of the compound's formula
- iii. Percent composition is the relative volume of each element in the compound
- iv. Two compounds can have identical percentage compositions

- A) i, ii B) i, ii, iii C) ii, iii D) i, iv E) iv only

7. Which of the following is a candidate for an empirical formula?

- A) $\text{C}_{24}\text{H}_{40}\text{O}_{22}$ B) N_2O_4 C) CH_3OH D) C_6H_6 E) $\text{Na}_2\text{S}_2\text{O}_8$

8. The mass percent of nitrogen, in nitric acid, HNO_3 , is

- A) 1.6% B) 11.1% C) 15.6% D) 22.2%

9. Hydrogen peroxide is known to contain 5.9% hydrogen and 94.1% oxygen, by mass. If its molar mass is 34.0 g/mol, its molecular formula is

- A) HO B) H_2O_2 C) HO_2 D) HNO_2

Grade 11 Stoichiometry Test

Kagendo Mutua, Cynthia S Sunal



Grade 11 Stoichiometry Test:

Grade 11 Stoichiometry Test Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Grade 11 Stoichiometry Test**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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