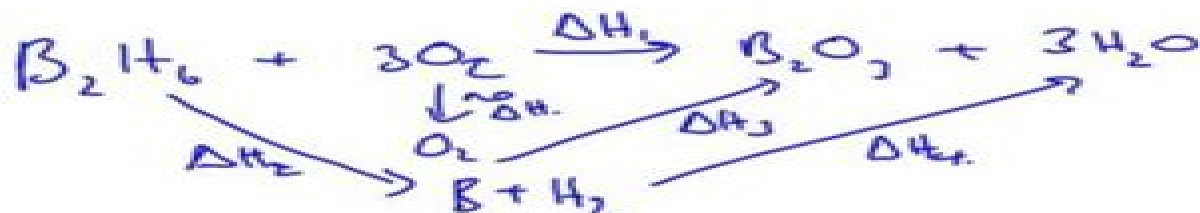




Calculate the enthalpy of combustion of diborane (B_2H_6) in $kJ\ mol^{-1}$, using the following data.

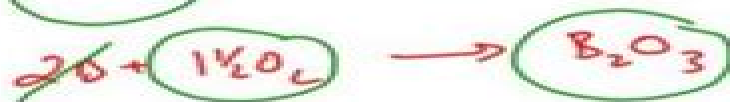
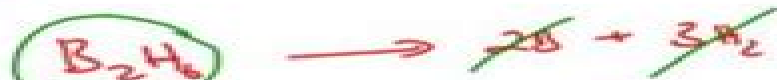


By Hess's Law $\Delta H_1 = \Delta H_2 + \Delta H_3 + \Delta H_4$.

Equations to use.



Calculation



$$\Delta H_1 = \Delta H_2 + \Delta H_3 + \Delta H_4 = -36 - 1274 - 858 = -2168 \text{ kJ mol}^{-1}$$

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Zhihan Nan, Sheng Zhang



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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Table of Contents Hess Law Problems With Answers

1. Understanding the eBook Hess Law Problems With Answers
 - The Rise of Digital Reading Hess Law Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Hess Law Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hess Law Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hess Law Problems With Answers
 - Personalized Recommendations
 - Hess Law Problems With Answers User Reviews and Ratings
 - Hess Law Problems With Answers and Bestseller Lists
5. Accessing Hess Law Problems With Answers Free and Paid eBooks

- Hess Law Problems With Answers Public Domain eBooks
 - Hess Law Problems With Answers eBook Subscription Services
 - Hess Law Problems With Answers Budget-Friendly Options
6. Navigating Hess Law Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Hess Law Problems With Answers Compatibility with Devices
 - Hess Law Problems With Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hess Law Problems With Answers
 - Highlighting and Note-Taking Hess Law Problems With Answers
 - Interactive Elements Hess Law Problems With Answers
 8. Staying Engaged with Hess Law Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hess Law Problems With Answers
 9. Balancing eBooks and Physical Books Hess Law Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hess Law Problems With Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Hess Law Problems With Answers
 - Setting Reading Goals Hess Law Problems With Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Hess Law Problems With Answers
 - Fact-Checking eBook Content of Hess Law Problems With Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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