

# Acids & Bases Study Guide

6. Water is considered neutral. Describe what this means in terms of  $[H^+]$  &  $[OH^-]$  concentrations.

8. Complete the following:

$$pH = -\log[H^+]$$

$$K_w = [H^+][OH^-]$$

7. Fill in the blank:

9. What is the  $[OH^-]$  of a solution with a pH of 10?

10. What is the  $[H^+]$  of a solution with a pH of 4?

11. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-5}$  M?

12. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-4}$  M?

13. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-12}$  M?

14. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-10}$  M?

15. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-11}$  M?

16. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-13}$  M?

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

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1. Fill in the blanks or complete and contrast the properties of Acids & Bases.

| Acids                                                                                                                                                                               | Bases                                                                                                                                                                               |
|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <ul style="list-style-type: none"> <li>• Corrosive _____</li> <li>• Tastes _____</li> <li>• Feels _____</li> <li>• Turns litmus paper _____</li> <li>• Reacts with _____</li> </ul> | <ul style="list-style-type: none"> <li>• Corrosive _____</li> <li>• Tastes _____</li> <li>• Feels _____</li> <li>• Turns litmus paper _____</li> <li>• Reacts with _____</li> </ul> |

2. Fill in the chart for the following acids & bases.

| Acid or Base? | Strong or Weak? | Formula    | Name              |
|---------------|-----------------|------------|-------------------|
|               |                 |            | hydrochloric acid |
|               |                 |            | sulfuric acid     |
|               |                 | $HNO_3$    |                   |
|               |                 | $NaOH$     |                   |
|               |                 | $Ca(OH)_2$ | hydroxide acid    |

3. Predict the products for the following double replacement reactions involving acids & bases. Identify which component is the Acid, Base, Conjugate Acid (CA), and Conjugate Base (CB) in the formula.



4. Label the pH Scale from 0 to 14. Acidic, Basic, Neutral, Strongest Acid (Strongest Base, Weak Acid, Weak Base).



# Holt Chemistry Acids And Bases Study Guide

**Christian Drosten**



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