

# **MCAT Chemistry Equation Sheet**

#### ORBITALS AND ELECTRONS

Quantum numbers descriptions

Quantum mamber	Description	Possible values
Principal quantum manber (n)	energy level	n=1, 2, 3
Angular monocotum quantum number (f)	orbital shape (s, p, d, f)	J = 0, sos- 1
Magnetic quantum number (m)	orbital oriomation (p., p., p.)	and the state of
Spile quantum number (m.)	describes spin of electron	100, 11 11 115 OK

Each electron has a unique set of quantum numbers.

#### ELECTRONIC CONFIGURATION

Electron configuration: describes the location of all of the electrons in an element, e.g. Carbon = ta/2a/2bc/

Heisensberg Uncertainty Principle: the position and monopolism of a particle cannot both be exactly known at the same time. Pauli exclusion principle: no two electrons in an attorn can have the name set of four quantum numbers (n, l, m, m,) which means that each orbital with quantum numbers (n, l, m) can hold at most two electrons with reposition seins.

Bland's rule: when electrons added to orbitals of equal energy, a single electron enters each orbital before a second electron with apposite spin cuters any orbital. The spins remain parallel if possible, e.g.

Notrogem   N = T	74	7.4	1.1
T.	24	24	2p., 2p.,

Author Principle: dictates the order in which orbitals fill. Some exceptions occur in the transition elements.

14 2s: 2p 3s: 3p 3d ds: 4p 4d 4f 5s: 5p 5d 5f ds: 6p 6d 6f

#### THE GAS PHASE

#### Units of Pressure:

Table = 760 rhmHg = 760 torr = 101.325 kPa = 1.01325 barr

**CONTRACTORS WORTHWATERS** 

- molecules are very small compared to the distance b'w them.
- 2. molocules are continuously moving
- pressure of the gas collisions of the meteories with container walls
- 4. molecules do not experience intermolecular

Ideal-gas law: PV = nRT

Combined gas leve:  $\frac{P_i V_i}{T_i} \approx \frac{P_i V_j}{T_i}$ 

Charles' law:  $\frac{V_{i}}{T_{i}} = \frac{V_{i}}{T_{i}}$   $P = \frac{V_{i}}{T_{i}}$ 

Bestr's law:  $PE = P_1F_2$   $F = \phi_0$ 

No one,  $\Gamma$  Law:  $\frac{P_1}{r} = \frac{P_2}{r}$   $\Gamma$  = cons

STP – standard temperature (0°C or 273.15 K) and provious (1 atm)

I mol of ideal gas occupies F = 22.4 L standard conditions - 25°C and 1 atm

Dialton's law of partial pressure (misture of the gases)

$$P = p_x + p_x + ... + p_x$$
  
 $p_x = \gamma_x \cdot P$ 

$$\chi_x = \frac{n_x}{n_{\text{max}}} = \frac{N_x}{N_{\text{max}}}$$

GRAHAM'S LAW OF EFFUSION

$$\frac{v_s}{v_s} = \sqrt{\frac{m_s}{m_s}}$$

$$KE_A = \frac{1}{2}m_A v_A^2$$

$$KE_A = KE_B = T_A = T_B$$
 bic  $KE \propto T$  , but

$$v_A \neq v_B \neq T_A = T_B$$

What leads to the pas "idealty", high T & low P & lower molecular weight or small size of gas molecules

#### Real cases:

- have intermolecular attractions (non-elastic optisions) called Van der Waals forces.
- contain molecules that have volume (or mean ratio site)

$$\left(P_{net} + a\frac{n^2}{V^2}\right) \cdot \left(V_{net} - nb\right) = nRT$$

#### INTERMOLECULAR FORCES

#### DISPERSION FORCES:

Describe the attraction between the negatively charged electron cloud of one molecule and the positively charged nuclei of neighbouring molecules. Descuar Forces:

Describe the attraction between the negatively charged end of a polar molecule and the positively charged ends of neighbouring polar molecules Hydeocen Bosones:

Inviolves fore pairs of electrons on an electronegative atom of one molecule and a polar bond to hydrogen in another molecule, they are confined to molecules that contain O, N, and F stores.

#### Procamicanicity:

The ease with which the electron clouds are distorted

#### MOUROULAN SOURCE

Molecules in molecular solids are held in place by the types of forces: dispersion forces, dipolar interactions, and/or hydrogen bonds

#### METALLIC SOLICE

Atoms in metallic solids are held in place by

delocalised bonding

#### Name of the Name of Street, Spirit, Sp

Contains an array of covalent bonds linking every atom to its neighbours

#### Joseph Soulo:

Contains cations and anions, attracted to one another by coulombic interactions

#### SOLUTIONS

molarity (M)	modes of solute/volume of solution. M=s/V
modulating days	moles of solute/kg of solvent
mode fraction (g)	medies of solute/total motes
mass percent	mass of solute x 190/ total mass of solution
parts per million (ppm)	mass of solute s 10°/total mass of solution

#### Repult's Law (distillation)

$$P_x = X_x P_x^{\alpha}$$

Where  $P_d$  = partial pressure of A in solution  $X_d$  = mode fraction of A in the liquid  $P_d^*$  = vapour pressure of pure A

# Colligative properties (melting/boiling points etc.)

Frocering point depression; salt on the sidewalk in winter lowers the focuring point of H<sub>2</sub>O to prevent the water from froceing

$$\Delta T_{\nu} = 3c_{\nu} m_{\nu}$$

Booling point elevation: salt in pasts water increases the boiling point of H<sub>2</sub>O so that the water boils hotter (and year pasts cooks more queckly).

$$\Delta T_{\rm in} = R_{\rm in} m T$$

Osimotic pressure: increase in pressure due to a solvent crossing a membrane into a mosty concentrated solution. ITV—mRT



# **Mcat Chemistry Review Sheet**

**Thomas Griffiths** 

**Mcat Chemistry Review Sheet:** 

# Mcat Chemistry Review Sheet Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "Mcat Chemistry Review Sheet," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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