

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
RESEARCH REPORT

RESEARCH REPORT

1. **Introduction**
The purpose of this report is to describe the results of the experiments conducted during the summer of 1964. The experiments were designed to study the effect of temperature on the rate of reaction between hydrogen peroxide and potassium permanganate in the presence of ceric sulfate as a catalyst. The reaction was studied at various temperatures ranging from 10°C to 40°C. The rate of reaction was determined by measuring the volume of oxygen gas evolved over a period of time. The results of the experiments are presented in the following sections.

2. **Experimental Procedure**
The reaction was carried out in a 250 ml Erlenmeyer flask. A known volume of 0.1 M potassium permanganate solution was added to a known volume of 0.1 M ceric sulfate solution. A known volume of 3% hydrogen peroxide solution was then added to the mixture. The flask was stoppered and placed in a water bath at the desired temperature. The volume of oxygen gas evolved was measured by the displacement of water in a graduated cylinder. The rate of reaction was calculated from the volume of oxygen gas evolved per unit time.

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