

Gas Stoichiometry Exam
2016-2017

Name KLAMY KEY
Period _____ Date _____

Multiple Choice

Refer to the set of graphs below when answering the next two questions:

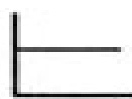
Graph #1



Graph #2



Graph #3



Graph #4



- B 1. Which of the graphs above is the best qualitative graph of pressure versus volume of nitrogen at constant temperature and number of moles?

A) Graph #1
B) Graph #2
C) Graph #3
D) Graph #4

→ INVERSE

- A 2. Which of the graphs above is the best qualitative graph of Kelvin temperature versus pressure of an ideal gas at constant volume and number of moles?

A) Graph #1
B) Graph #2
C) Graph #3
D) Graph #4

→ DIRECT

- D 3. Use the kinetic molecular theory of gases to predict what would happen to a closed sample of a gas whose temperature increased while its volume decreased.

→ MOVE FASTER, MORE COLLISIONS
A) The average kinetic energy of the molecules of the gas would decrease.
B) Its pressure would decrease.
C) The number of moles of gas would decrease.
D) Its pressure would increase.

- B 4. A hydrocarbon gas with the empirical formula CH_2 has a density of 1.3 g/L at 0°C and 1.00 atm. A possible formula for the hydrocarbon is:

A) CH_2
B) C_2H_4
C) C_3H_6
D) C_4H_8
E) C_5H_{10}

~~1.3 g/L~~

$$1.3 \text{ g/L} \times 22.4 \text{ L/mol} \approx 30 \text{ g/mol}$$

Holt Modern Chemistry Gases Test A Answers

Harold Clark Metcalfe



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1. Understanding the eBook Holt Modern Chemistry Gases Test A Answers
 - The Rise of Digital Reading Holt Modern Chemistry Gases Test A Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Modern Chemistry Gases Test A Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Modern Chemistry Gases Test A Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Modern Chemistry Gases Test A Answers
 - Personalized Recommendations
 - Holt Modern Chemistry Gases Test A Answers User Reviews and Ratings
 - Holt Modern Chemistry Gases Test A Answers and Bestseller Lists
5. Accessing Holt Modern Chemistry Gases Test A Answers Free and Paid eBooks
 - Holt Modern Chemistry Gases Test A Answers Public Domain eBooks
 - Holt Modern Chemistry Gases Test A Answers eBook Subscription Services
 - Holt Modern Chemistry Gases Test A Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
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 - Holt Modern Chemistry Gases Test A Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Modern Chemistry Gases Test A Answers
 - Highlighting and Note-Taking Holt Modern Chemistry Gases Test A Answers
 - Interactive Elements Holt Modern Chemistry Gases Test A Answers
8. Staying Engaged with Holt Modern Chemistry Gases Test A Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Modern Chemistry Gases Test A Answers
9. Balancing eBooks and Physical Books Holt Modern Chemistry Gases Test A Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Modern Chemistry Gases Test A Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Modern Chemistry Gases Test A Answers
 - Setting Reading Goals Holt Modern Chemistry Gases Test A Answers
 - Carving Out Dedicated Reading Time
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 - Distinguishing Credible Sources
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