

MODUL BELAJAR KIMIA 8

Kompetensi Dasar : 1.1	Memahami Struktur atom,Sifat-sifat Unsur dan Ikatan kimia	Periodik
Materi Pokok :	1. Ikatan kovalen Polar dan Non Polar 2. Ikatan Logam	
Waktu :	2 x 45 menit	

1. IKATAN KOVALEN POLAR

Atom-atom pembentuknya mempunyai gaya tarik yang tidak sama terhadap pasangan elektron persekutuan. Hal ini terjadi karena beda keelektronegatifan kedua atomnya. Elektron persekutuan akan bergeser ke arah atom yang lebih elektronegatif akibatnya terjadi pemisahan kutub positif dan negatif. Dalam senyawa HCl ini, Cl mempunyai keelektronegatifan yang lebih besar dari H, sehingga pasangan elektron lebih tertarik ke arah Cl, akibatnya H relatif lebih elektropositif sedangkan Cl relatif menjadi elektronegatif.

Pemisahan muatan ini menjadikan molekul itu bersifat polar dan memiliki "momen dipol" sebesar:

$$T = q \cdot l$$

dimana :

T = momen dipol

q = kelebihan muatan pada masing-masing atom

l = jarak antara kedua inti atom

2. IKATAN KOVALEN NON POLAR

Tidak muatan negatif elektron persekutuan berhimpit, sehingga pada molekul pembentuknya tidak terjadi momen dipol, dengan perkataan lain bahwa elektron persekutuan mendapat gaya tarik yang sama. Contoh: Kedua atom H mempunyai harga keelektronegatifan yang sama. Karena arah tarikan simetris, maka tidak muatan negatif elektron persekutuan berhimpit. Contoh lain adalah senyawa CO_2 , O_2 , Br_2 dan lain-lain

LATIHAN

1. Di antara molekul berikut ini tentukan yang ikatannya polar dan yang non polar

a. CO_2

c. F_2

b. Br_2

d. PCl_3

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CL Gary



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Table of Contents Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1

1. Understanding the eBook Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - The Rise of Digital Reading Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Advantages of eBooks Over Traditional Books
2. Identifying Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Personalized Recommendations
 - Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 User Reviews and Ratings
 - Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 and Bestseller Lists
5. Accessing Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 Free and Paid eBooks

- Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 Public Domain eBooks
- Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 eBook Subscription Services
- Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 Budget-Friendly Options
- 6. Navigating Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 Compatibility with Devices
 - Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Highlighting and Note-Taking Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Interactive Elements Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
- 8. Staying Engaged with Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
- 9. Balancing eBooks and Physical Books Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Setting Reading Goals Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Fact-Checking eBook Content of Kharisma Smk Teknik Modul Kimia Kelas 12 Smt1
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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