



CHEMISTRY

CHEAT SHEET

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EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol $\rightleftharpoons$. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
Equilibrium Constant, K_c For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant, K_p For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
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Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration	Increasing concentration of a reactant will shift the equilibrium to the right. Increasing concentration of a product will shift the equilibrium to the left.	For a general reaction: $aA + bB \rightleftharpoons cC + dD$ Increasing concentration of A or B will shift the equilibrium to the right, increasing the concentration of C and D.
Pressure	For a general reaction, if the number of moles of gas is different on the two sides of the equation, then the equilibrium will shift to the side with the fewer moles of gas.	For a general reaction: $aA + bB \rightleftharpoons cC + dD$ If the number of moles of gas is different on the two sides of the equation, then the equilibrium will shift to the side with the fewer moles of gas.
Temperature	For an exothermic reaction, increasing temperature will shift the equilibrium to the left. For an endothermic reaction, increasing temperature will shift the equilibrium to the right.	For an exothermic reaction: $aA + bB \rightleftharpoons cC + dD$ Increasing temperature will shift the equilibrium to the left, decreasing the concentration of C and D.

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