

Phases of Matter Vocabulary

Word	Definition
Avogadro's Hypothesis	Equal volumes of two ideal gases under the same conditions of temperature and pressure will contain equal number of molecules.
Boiling	The transition of a liquid into a gas at the boiling point.
Condensing	The transition of a gas into a liquid at the boiling point.
Deposition	The transition of a gas into a solid.
Equilibrium	The condition that exists when the rates of two opposing changes are equal.
Evaporating	The transition of the surface molecules of a liquid into a gas below the boiling point.
Freezing	The transition of a liquid into a solid at the freezing point.
Gas	A phase of matter characterized by the complete dissociation of matter particles from each other with the distances between the particles very large in comparison to the size of the particles and no attractive forces between them.
Heat of Fusion	The energy required to melt a gram of solid at its melting point.
Heat of Vaporization	The energy required to boil a gram of liquid at its boiling point.
Ideal Gas	A gas in which the molecules are infinitely small and far apart, the molecules travel with a straight-line motion, all collisions have no net loss of energy (elastic), there are no attractive forces between molecules and the speed of the molecules is directly proportional to the Kelvin temperature. Gases are most ideal at high temperature and low pressure.
Liquid	A phase of matter characterized by matter loosely organized yet kept together by intermolecular or ionic attractive forces.
Melting	The transition of a solid into a liquid at the melting point.
Pressure	Force exerted over an area.
Solid	A phase of matter characterized by matter arranged in regular geometric patterns called crystal lattices with only vibration motion, no relative motion.
Specific Heat	The energy required to heat one gram of a substance by one Kelvin.
Sublimation	The transition of a solid into a gas.
Vapor pressure	The pressure exerted by vapor in a vapor-liquid mixture in a closed system at equilibrium.
Vapor-Liquid Equilibrium	A system where the rate of evaporation equals the rate of condensing.

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