

Organic Chemistry

Word	Definition
Addition reaction	Halogen atoms break the double or triple bond on an unsaturated hydrocarbon chain and bond to either side of where the bond was broken.
Alcohol	A hydrocarbon with an $-OH$ (hydroxyl) group somewhere on the chain.
Aldehyde	A hydrocarbon with a $-CO$ (carbonyl) group bonded onto a primary carbon.
Alkane	A hydrocarbon with the general formula C_nH_{2n+2} , where all of the carbon-to-carbon bonds are single bonds.
Alkene	A hydrocarbon with the general formula C_nH_{2n} , where one of the carbon-to-carbon bonds is a double bond.
Alkyl group	An alkane fragment substituted onto a hydrocarbon chain, example methyl.
Alkyne	A hydrocarbon with the general formula C_nH_{2n-2} , where one of the carbon-to-carbon bonds is a triple bond.
Allotrope	A molecular form of an element. Oxygen has two allotropes: O_2 (diatomic oxygen) and O_3 (ozone).
Amide	A hydrocarbon with a $-CO-NH-$ (amide) group somewhere on the chain.
Amine	A hydrocarbon with a $-NH-$ (amine) group somewhere on the chain.
Combustion	Organic molecule reacts in the presence of oxygen to form carbon dioxide and water vapor. This is a highly exothermic reaction also known as "burning".
Dehydration synthesis	The joining of two organic molecules by the removal of water by a catalyst (dehydrating agent), often concentrated sulfuric acid).
Ester	A hydrocarbon with a $-COO$ (carboxyl) group bonded onto a secondary carbon.
Esterification	The dehydration synthesis of an ester by reacting an organic acid with a primary alcohol.
Ether	A molecule consisting of two alkyl groups on either side of an oxygen atom.
Etherification	The dehydration synthesis of an ether by reacting two molecules of primary alcohol.
Fermentation	The anaerobic digestion of simple sugars by yeast to produce ethanol and carbon dioxide.
Halocarbon	A hydrocarbon that has had one or more halogen atoms substituted or added to the carbon chain.
Hydrocarbon	An organic molecule composed of hydrogen and carbon.
Isomer	Molecules with the same molecular formula, but with different structural formulas
Ketone	A hydrocarbon with a $-CO$ (carbonyl) group bonded onto a secondary carbon.
Monomer	A single molecule, usually an alkene or alkydiene (addition polymerization) or a diol and dicarboxylic acid (dehydration polymerization).
Organic acid	A hydrocarbon with a $-COOH$ (carboxyl) group bonded onto a primary carbon.
Polymer	A huge chain of connected monomers. Examples include DNA, rayon, silk, polybutadiene and polyisoprene (rubber), polypropylene, polyvinyl chloride, polytetrafluoroethylene, polystyrene (plastic)
Polymerization	The joining of monomer units by addition reactions or dehydration synthesis to form enormous macromolecules called polymers.
Primary	Positional description of a carbon on the end of a hydrocarbon chain that is only directly bonded to one other carbon atom.
Saponification	The hydrolysis of a glycerol ester (fat) by a strong base to form glycerol and soap.
Saturated hydrocarbon	A hydrocarbon with all single carbon-carbon bonds.
Secondary	Positional description of a carbon in the middle of a hydrocarbon chain that is directly bonded to two other carbon atoms.
Substitution reaction	Halogen atoms replace hydrogen atoms on a saturated hydrocarbon chain.
Tertiary	Positional description of a carbon in the middle of a hydrocarbon chain that is directly bonded to three other carbon atoms.
Unsaturated Hydrocarbon	A hydrocarbon with one or more double or triple carbon-carbon bond.

Mark Rosengarten Organic Chemistry Answers

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