

Name _____

Gas Laws

21. Q: What is the relationship between volume and moles in Avogadro's Principle?

A: Volume is directly proportional to the number of moles.

22. Q: Calculate the pressure of 1.0 mole of gas in a 10.0 L container at 273 K using the ideal gas law.

A: The pressure is 2.24 atm.

23. Q: According to the ideal gas law, what happens to the pressure if the volume is reduced by half while the temperature and number of moles remain constant?

A: The pressure doubles.

24. Q: If the volume of a gas is tripled and the temperature is doubled, what happens to the pressure according to the combined gas law?

A: The pressure decreases to 2/3 of its original value.

25. Q: Using gas stoichiometry, how many liters of CO_2 are produced from 5.0 L of O_2 in the reaction $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ at STP?

A: 5.0 L of CO_2 are produced.

26. Q: In the reaction $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$, how many liters of NO_2 are produced from 4.0 L of NO at STP?

A: 4.0 L of NO_2 are produced.

27. Q: Calculate the volume of N_2 gas required to produce 20.0 L of NH_3 in the reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ at STP.

A: 10.0 L of N_2 are required.

28. Q: If the temperature of a gas is increased from 300 K to 600 K and the volume is kept constant, what happens to the pressure?

A: The pressure doubles.

29. Q: Using the ideal gas law, how do you calculate the volume of gas if you know the pressure, temperature, and number of moles?

A: $V = nRT/P$

30. Q: In the reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$, how many liters of HCl are produced from 3.0 L of H_2 at STP?

A: 6.0 L of HCl are produced.



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Thomas Griffiths



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