

**Problem Solving** *continued*

- 12.** Urea,  $(\text{NH}_2)_2\text{CO}$ , is an important fertilizer that is manufactured by the following reaction.



What volume of  $\text{NH}_3$  at STP will be needed to produce  $8.50 \times 10^4$  kg of urea if there is an 89.5% yield in the process?

- 13.** An obsolete method of generating oxygen in the laboratory involves the decomposition of barium peroxide by the following equation.



What mass of  $\text{BaO}_2$  reacted if 265 mL of  $\text{O}_2$  is collected by water displacement at 0.975 atm and  $10.^\circ\text{C}$ ?

- 14.** It is possible to generate chlorine gas by dripping concentrated  $\text{HCl}$  solution onto solid potassium permanganate according to the following equation.



If excess  $\text{HCl}$  is dripped onto 15.0 g of  $\text{KMnO}_4$ , what volume of  $\text{Cl}_2$  will be produced? The  $\text{Cl}_2$  is measured at  $15^\circ\text{C}$  and 0.959 atm.

- 15.** Ammonia can be oxidized in the presence of a platinum catalyst according to the following equation.



The  $\text{NO}$  that is produced reacts almost immediately with additional oxygen according to the following equation.



If 35.0 kL of oxygen at STP react in the first reaction, what volume of  $\text{NH}_3$  at STP reacts with it? What volume of  $\text{NO}_2$  at STP will be formed in the second reaction, assuming there is excess oxygen that was not used up in the first reaction?

- 16.** Oxygen can be generated in the laboratory by heating potassium chlorate. The reaction is represented by the following equation.



What mass of  $\text{KClO}_3$  must be used in order to generate 5.00 L of  $\text{O}_2$ , measured at STP?

- 17.** One of the reactions in the Solvay process is used to make sodium hydrogen carbonate. It occurs when carbon dioxide and ammonia are passed through concentrated salt brine. The following equation represents the reaction.



- a.** What volume of  $\text{NH}_3$  at  $25^\circ\text{C}$  and 1.00 atm pressure will be required if 38 000 L of  $\text{CO}_2$ , measured under the same conditions, react to form  $\text{NaHCO}_3$ ?
- b.** What mass of  $\text{NaHCO}_3$  can be formed when the gases in (a) react with  $\text{NaCl}$ ?

# Holt Chemistry Stoichiometry Problem Solving

**Richard Bailey**



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