

Chapter 9 Study Guide

Holt Chemistry: Stoichiometry

Questions

1. What is stoichiometry? Finding mass of product formed (or reactant consumed) when given a mass of another species in the reaction, by using the mole to mole ratio from the balanced equation.
2. A balanced equation verifies the law of conservation of mass: Total mass of reactants must equal total mass of products.
3. Give the 3 steps of converting the mass (grams) of one substance to the mass (grams) of another substance: 1. Convert grams to moles using molar mass; 2. Mole to Mole ratio from balanced equation; 3. Convert moles back to grams using molar mass

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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 *ChemDiscovery Teacher Edition* Olga I. Agapova, 2002 **General Chemistry: Experiment and Theory** Richard H. Eastman, 1970 **International Chemistry Directory**, 1969 Intended as a comprehensive current source of professional information for the use of chemists and biochemists Main body of book is Academic departments and faculties alphabetically arranged by name of the institution in which chairmen and faculty of chemistry departments are identified Laboratories societies meetings grants fellowships graduate support awards books and journals also included in separate sections Faculty name index Technology in the Curriculum: Science resource guide, 1986 **Project SERAPHIM 1991 Catalog**, 1991 **AAAS Science Book List, 1978-1986** Kathryn Wolff, Susan M. O'Connell, Valerie J. Montenegro, 1986 A selected and annotated list of science and mathematics books which supplements the AAAS science book list 3rd ed 1970 and the AAAS science book list supplement 1978 **The Software Encyclopedia**, 1988 *Analytical Chemistry* Douglas A. Skoog, Donald M. West, 1979

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