



ANSWER GUIDE
PHYSICAL SCIENCES: GRADE 12
CLASS TEST: ACIDS & BASES

Written on: 25.06.2020

Total Marks: 65

1	1.1	B ✓✓	2
	1.2	C ✓✓	2
2	Hydrochloric acid is a strong acid and sodium hydroxide is a strong base, ✓ therefore the equivalence point will be reached at a pH close to 7, ✓ which is in the colour change range of bromothymol blue.		2
3	3.1	An acid is a proton donor. ✓	1
	3.2	A concentrated acid is an acid solution with a high proportion of acid (solute) to water (solvent). ✓✓ regardless of the degree of ionisation or dissociation of the acid in the solution. A strong acid is an acid that dissociates or ionises almost completely when in solution, regardless of its concentration. ✓✓	4
4	4.1	LiOH ✓	5
	4.2	H ₃ PO ₄ ✓	
	4.3	HNO ₃ ✓	
	4.4	HCOOH ✓	
	4.5	NaHCO ₃ ✓	
5	$\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} \checkmark \rightarrow 2\text{H}_3\text{O}^+ \checkmark + \text{SO}_4^{2-} \checkmark$ OR: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$		3
6	6.1	A standard solution is a solution of known concentration. ✓✓	2
	6.2	$c = \frac{m}{MV} \quad \therefore m = cMV \checkmark$ $= (0,15)(106)(0,25) \checkmark$ $= \underline{3,975 \text{ g Na}_2\text{CO}_3} \checkmark$	3
	6.3	$\frac{n_a}{n_b} = \frac{c_a V_a}{c_b V_b} \quad \therefore c_a = \frac{n_a c_b V_b}{n_b V_a} \checkmark$ $= \frac{(1)(0,15)(25)}{(1)(27,5)} \checkmark$ $= \underline{1,38 \times 10^{-1} \text{ mol.dm}^{-3}} \checkmark$	4
7	pH = $-\log[\text{H}^+] \checkmark$ and: HCl dissociates completely $\therefore 0,2 \text{ mol.dm}^{-3} \text{ H}^+$ ions = $-\log[0,2] \checkmark$ = $0,69 \checkmark$		3

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**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**



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molecular orbital MO theory of bonding galvanic cells and chemical thermodynamics are considered next Separate chapters are devoted to MO theory of covalent and metallic bonding orbital hybridization intermolecular forces acids and bases ionic equilibrium calculations and polymers and biochemicals This monograph is intended for chemistry students Molecular World Catherine M. Jackson, 2023-07-04 A compelling and innovative account that reshapes our view of nineteenth century chemistry explaining a critical period in chemistry's quest to understand and manipulate organic nature According to existing histories theory drove chemistry's remarkable nineteenth century development In Molecular World Catherine M Jackson shows instead how novel experimental approaches combined with what she calls laboratory reasoning enabled chemists to bridge wet chemistry and abstract concepts and in so doing create the molecular world Jackson introduces a series of practice based breakthroughs that include chemistry's move into lampworked glassware the field's turn to synthesis and subsequent struggles to characterize and differentiate the products of synthesis and the gradual development of institutional chemical laboratories an advance accelerated by synthesis and the dangers it introduced Jackson's historical reassessment emerges from the investigation of alkaloids by German chemists Justus Liebig August Wilhelm Hofmann and Albert Ladenburg Stymied in his own research Liebig steered his student Hofmann into pioneering synthesis as a new investigative method Hofmann's practice based laboratory reasoning produced a major theoretical advance but he failed to make alkaloids That landmark fell to Ladenburg who turned to cutting edge theory only after his successful synthesis In telling the story of these scientists and their peers Jackson reveals organic synthesis as the ground chemists stood upon to forge a new relationship between experiment and theory with far reaching consequences for chemistry as a discipline **Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb** Wade A. Freeman, 1986

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web jun 3 2021 issn online 2052 2819 references the calculation of fluid balance involves measuring the amount of fluid

entering the body and comparing it to the amount of fluid leaving the body the purpose is to determine whether there is a deficit or an excess of fluid bannerman 2018

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